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WORKSHEET ON SINGLE & DOUBLE REPLACEMENT REACTIONS

Predict the products. Write formulas & balance each reaction. If there is no reaction, then just put NO RXN.

Single Replacement: $A + BC \rightarrow B + AC$ or $A + BC \rightarrow C + BA$ (when A and C are negative ions)

1. Zinc + Hydrogen chloride $\rightarrow$
2. Magnesium + Hydrogen Sulfate $\rightarrow$
3. Copper (II) chloride + Fluorine $\rightarrow$
4. Silver + Sodium Hydroxide $\rightarrow$
5. Potassium iodide + Bromine $\rightarrow$
6. Calcium + Hydrogen hydroxide $\rightarrow$
7. Iron IV oxide + Hydrogen $\rightarrow$

Double Replacement: $AB + CD \rightarrow AD + CB$

1. Barium chloride + Aluminum sulfate $\rightarrow$
2. Calcium nitride + water $\rightarrow$
3. Calcium hydroxide + Hydrogen phosphate $\rightarrow$
4. Hydrogen sulfate + Sodium hydrogen carbonate $\rightarrow$
5. Calcium hydroxide + Ammonium chloride $\rightarrow$
6. Potassium iodide + Lead II Nitrate $\rightarrow$
7. Sodium acetate + Calcium sulfide $\rightarrow$

Complete each word equation, write formulas and balance the reaction equation. Then identify and place the type of reaction (single replacement or double replacement) in the blank provided.

1. Zinc + Silver nitrate $\rightarrow$
2. Aluminum + Hydrogen chloride $\rightarrow$
3. Magnesium oxalate + Ammonium carbonate $\rightarrow$
4. Calcium + Aluminum nitrate $\rightarrow$
5. Potassium fluoride + Lead (II) Nitrate $\rightarrow$
6. Calcium bromide + Silver nitrate $\rightarrow$
7. Ammonium phosphate + Barium acetate $\rightarrow$
8. Sodium chloride + Potassium $\rightarrow$
9. Magnesium nitrate + ammonium chloride $\rightarrow$
10. Iron (III) chloride + calcium $\rightarrow$
11. Chlorine + Sodium bromide $\rightarrow$
12. Potassium chloride + Silver nitrate $\rightarrow$
13. Calcium hydroxide + Hydrogen nitrate $\rightarrow$
14. Lead II nitrate + Potassium chloride $\rightarrow$
15. Strontium carbonate + Hydrogen nitrate $\rightarrow$
16. Gold + Potassium nitrate $\rightarrow$
17. Zinc + Silver nitrate $\rightarrow$
18. Aluminum + Copper II sulfate $\rightarrow$

Element	Symbol	Atomic Number	Group	Period	Block
H	H	1	1	1	s
He	He	2	18	1	s
Li	Li	3	1	2	s
Be	Be	4	2	2	s
B	B	5	13	2	p
C	C	6	14	2	p
N	N	7	15	2	p
O	O	8	16	2	p
F	F	9	17	2	p
Ne	Ne	10	18	2	p
Na	Na	11	1	3	s
Mg	Mg	12	2	3	s
Al	Al	13	13	3	p
Si	Si	14	14	3	p
P	P	15	15	3	p
S	S	16	16	3	p
Cl	Cl	17	17	3	p
Ar	Ar	18	18	3	p
K	K	19	1	4	s
Ca	Ca	20	2	4	s
Sc	Sc	21	3	4	d
Ti	Ti	22	4	4	d
V	V	23	5	4	d
Cr	Cr	24	6	4	d
Mn	Mn	25	7	4	d
Fe	Fe	26	8	4	d
Co	Co	27	9	4	d
Ni	Ni	28	10	4	d
Cu	Cu	29	11	4	d
Zn	Zn	30	12	4	d
Ga	Ga	31	13	4	p
Ge	Ge	32	14	4	p
As	As	33	15	4	p
Se	Se	34	16	4	p
Br	Br	35	17	4	p
Kr	Kr	36	18	4	p
Rb	Rb	37	1	5	s
Sr	Sr	38	2	5	s
Y	Y	39	3	5	d
Zr	Zr	40	4	5	d
Nb	Nb	41	5	5	d
Mo	Mo	42	6	5	d
Tc	Tc	43	7	5	d
Ru	Ru	44	8	5	d
Rh	Rh	45	9	5	d
Pd	Pd	46	10	5	d
Ag	Ag	47	11	5	d
Cd	Cd	48	12	5	d
In	In	49	13	5	p
Sn	Sn	50	14	5	p
Sb	Sb	51	15	5	p
Te	Te	52	16	5	p
I	I	53	17	5	p
Xe	Xe	54	18	5	p
Ba	Ba	56	2	6	s
La	La	57	3	6	f
Ce	Ce	58	4	6	f
Pr	Pr	59	5	6	f
Nd	Nd	60	6	6	f
Pm	Pm	61	7	6	f
Sm	Sm	62	8	6	f
Eu	Eu	63	9	6	f
Gd	Gd	64	10	6	f
Tb	Tb	65	11	6	f
Dy	Dy	66	12	6	f
Ho	Ho	67	13	6	f
Er	Er	68	14	6	f
Tm	Tm	69	15	6	f
Yb	Yb	70	16	6	f
Lu	Lu	71	17	6	f
Hf	Hf	72	4	6	d
Ta	Ta	73	5	6	d
W	W	74	6	6	d
Re	Re	75	7	6	d
Os	Os	76	8	6	d
Ir	Ir	77	9	6	d
Pt	Pt	78	10	6	d
Au	Au	79	11	6	d
Hg	Hg	80	12	6	d
Tl	Tl	81	13	6	p
Pb	Pb	82	14	6	p
Bi	Bi	83	15	6	p
Po	Po	84	16	6	p
At	At	85	17	6	p
Fr	Fr	87	1	7	s
Ra	Ra	88	2	7	s
Ac	Ac	89	3	7	f
Th	Th	90	4	7	f
Pa	Pa	91	5	7	f
U	U	92	6	7	f
Np	Np	93	7	7	f
Pu	Pu	94	8	7	f
Am	Am	95	9	7	f
Cm	Cm	96	10	7	f
Bk	Bk	97	11	7	f
Cf	Cf	98	12	7	f
Es	Es	99	13	7	f
Fm	Fm	100	14	7	f
Md	Md	101	15	7	f
No	No	102	16	7	f
Lr	Lr	103	17	7	f

Balancing Equations Race

1. $C_2H_4 + O_2 \rightarrow CO_2 + H_2O$
2. $Al + Fe_2O_3 \rightarrow Al_2O_3 + Fe$
3. $Na + Cl_2 \rightarrow NaCl$
4. $H_2O_2 \rightarrow H_2O + O_2$
5. $C_2H_5OH + O_2 \rightarrow CO_2 + H_2O$
6. $H_2O + CO_2 \rightarrow C_2H_4 + O_2$
7. $NaCO_3 \rightarrow NaCl + O_2$
8. $(NH_4)_3PO_4 + Pb(NO_3)_2 \rightarrow Pb_3(PO_4)_2 + NH_4NO_3$
9. $BF_3 + Li_2SO_4 \rightarrow B_2(SO_4)_3 + LiF$
10. $CH_4 + O_2 \rightarrow CO_2 + H_2O$
11. $CaCO_3 + H_3PO_4 \rightarrow Ca_3(PO_4)_2 + H_2CO_3$
12. $Ag_2S \rightarrow Ag + S_8$
13. $KBr + Fe(OH)_3 \rightarrow KOH + FeBr_3$
14. $KNO_3 + H_2CO_3 \rightarrow K_2CO_3 + HNO_3$
15. $Pb(OH)_2 + CuO \rightarrow PbO_2 + CuOH$
16. $Cr(NO_3)_3 + (NH_4)_2SO_4 \rightarrow Cr_2(SO_4)_3 + NH_4NO_3$
17. $KOH + Ca_3(PO_4)_2 \rightarrow K_3PO_4 + Ca(OH)_2$
18. $Sn(NO_3)_2 + PbN_2 \rightarrow Sn_2N_4 + Pb(NO_3)_4$
19. $B_2Br_6 + HNO_3 \rightarrow B(NO_3)_3 + HBr$
20. $ZnS + AP \rightarrow ZnP_2 + AlS_3$

Chemical Bonding Worksheet

Write the Lewis structure for each molecule. Indicate the type of bonding that exists in the compound.

Compound	Electron Dot Structure	Electron Dot Structure	Bond Type
NO ₂			Covalent
NaCl			Ionic
FeS			Covalent
PCl ₃			Covalent
HgBr ₂			Covalent
CaO			Covalent
H ₂ O			Covalent
AgF			Covalent
Co-Br ₂			Covalent
Cl ₂			Covalent
CaS			Covalent
H ₂ S			Covalent
TeO ₂			Covalent
HF			Covalent
H ₂ Se			Covalent
H ₂ As ₂			Covalent
FeBr ₃			Covalent
PCl ₅			Covalent

ANSWER KEY

NO ₂	$\cdot \cdot \cdot \cdot$	$\cdot \cdot \cdot \cdot$	Covalent
NaCl	$\cdot \cdot \cdot \cdot$	$\cdot \cdot \cdot \cdot$	Ionic
FeS	$\cdot \cdot \cdot \cdot$	$\cdot \cdot \cdot \cdot$	Covalent
PCl ₃	$\cdot \cdot \cdot \cdot$	$\cdot \cdot \cdot \cdot$	Covalent
HgBr ₂	$\cdot \cdot \cdot \cdot$	$\cdot \cdot \cdot \cdot$	Covalent
CaO	$\cdot \cdot \cdot \cdot$	$\cdot \cdot \cdot \cdot$	Covalent
H ₂ O	$\cdot \cdot \cdot \cdot$	$\cdot \cdot \cdot \cdot$	Covalent
AgF	$\cdot \cdot \cdot \cdot$	$\cdot \cdot \cdot \cdot$	Covalent
Co-Br ₂	$\cdot \cdot \cdot \cdot$	$\cdot \cdot \cdot \cdot$	Covalent
Cl ₂	$\cdot \cdot \cdot \cdot$	$\cdot \cdot \cdot \cdot$	Covalent
CaS	$\cdot \cdot \cdot \cdot$	$\cdot \cdot \cdot \cdot$	Covalent
H ₂ S	$\cdot \cdot \cdot \cdot$	$\cdot \cdot \cdot \cdot$	Covalent
TeO ₂	$\cdot \cdot \cdot \cdot$	$\cdot \cdot \cdot \cdot$	Covalent
HF	$\cdot \cdot \cdot \cdot$	$\cdot \cdot \cdot \cdot$	Covalent
H ₂ Se	$\cdot \cdot \cdot \cdot$	$\cdot \cdot \cdot \cdot$	Covalent
H ₂ As ₂	$\cdot \cdot \cdot \cdot$	$\cdot \cdot \cdot \cdot$	Covalent
FeBr ₃	$\cdot \cdot \cdot \cdot$	$\cdot \cdot \cdot \cdot$	Covalent
PCl ₅	$\cdot \cdot \cdot \cdot$	$\cdot \cdot \cdot \cdot$	Covalent

Is this reaction to be balanced? Name: Benson Laird Date: 9/26/21 Student Exploration: Balance of the Qual Equations Knowledge Questions -Prior Questions (ask them before using the gizmo). The scouts are making customs of s What is bad in the image below? It is not balanced. 2. Assuming that a more than two graham cookies, a marshmallow and a piece of chocolate, how many customs s can be done with the ingredients shown? How many are in an oxygen molecule (O2)? 1. Excellent for an introduction based on consultation or a final task! 22. In the chemical equations of gizmo equilibrium, look at the floating cullets below the initial reaction: H2 + O2 à €™ H2O. A chemical equation is balanced if the number of each type of time on the left side is equal to the number of each type on the right side. Justify your answer the 5 pieces of Graham Graham cookies, but 2 cookies will only make one more. Each task will take approximately 1-2 class perities to complete, and each will provide students with an in-depth interpretation of their students about the balance of chemical reactions using this excellent detailed digital laboratory. How many are in a hydrogen molecule (H2)? This process is applied by a chemical equation. Justify your answer with evidence of the previous questions because there is an additional oxygen. 3 resolution: image size: download chemical reactions - bundleleeting unit to your students about the chemical reactions? This package will cover everything! This package contains eleven (11) detailed activities, work sheets, laboratories and evaluations that will guide your students through a complete unit of chemical reactions. Extremely à Ètil in a unit where physical ones can be differ from finding. This digital laboratory uses a PHET simulator to guide students through the following topics: identify reagents and products in a equation. Understanding the law of mass conservation in chemical equations. Balance equations using a visual representation of particles. Verification of results to verify that a chemical equation is properly balanced. This collection of worksheets has four (4) pages long and consists of: two (2) Questions/diagrams that involve the Phet SimulatorOne (1) Full equilibrium question page using the simulator to verify the results (1) Full review page with Balancesheet equations and guided instructions to help students learn how to use the simulator. Please wait while your URL is generated ... 23. In general, what is a subscript (such as the inamolecula of "2" in H2) on the molecule? (In the photo above). Perfect for students at all levels, and a perfect lesson for your chemical reaction unit. Each task will take about a class period to complete digital simulator laboratories: digital worksheet worksheets laboratories are the perfect tool to help students learn about complex chemical concepts. This package contains six (6) detailed worksheets to use with PHET simulators online. How many hydrogen and oxygen atoms are there in a water molecule (H2O)? 2H1O4. (H2O)? 2H1O4.

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